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Problem : Given that the free energy of formation of liquid water is  $-237 \text{ kJ / mol}$ , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate  $\Delta G$  for the reaction, which is  $-2 ( -237 \text{ kJ / mol} ) = 474 \text{ kJ / mol}$ . Knowing that  $\Delta G = -nFE^{\circ}$  and  $n$

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= 4, we calculate the potential is -1.23 V.

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Thermodynamics:-Problem 1:-A  
container holds a mixture of three  
nonreacting gases:  $n_1$  moles of the  
first gas with molar specific heat at  
constant volume  $C_{v1}$ , and so on. Find  
the molar specific heat at constant  
volume of the mixture, in terms of the  
molar specific heats

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Answers For Thermodynamics

Problems Answer for Problem # 1

Since the containers are insulated, no heat transfer occurs between the gas and the external environment, and since the gas expands freely into container B there is no resistance "pushing" against it, which means no work is done on the gas as it expands.

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problems and solutions. 1. 3000 J of  
heat is added to a system and 2500 J  
of work is done by the system. What is  
the change in internal energy of the  
system? Known : Heat ( $Q$ ) = +3000  
Joule. Work ( $W$ ) = +2500 Joule .  
Wanted: the change in internal energy  
of the system Solution :

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